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### GIET UNIVERSITY, GUNUPUR – 765022

M. Sc (First Semester) Examinations, May – 2021

## 20CHPC102- INORGANIC CHEMISTRY - I

(CHEMISTRY)

Time: 2 hrs Maximum: 50 Marks

#### (The figures in the right hand margin indicate marks.)

# PART – A

 $(2 \times 10 = 20 \text{ Marks})$ 

- Q.1. Answer ALL questions
  - a. Write the molecular orbital electronic configuration for HF molecule
  - b. What is meant by the term bond order? Calculate the bond order of  $N_2$ ,  $O_2$ ,  $O_2^+$  and  $O_2^-$
  - c. Give the differences between atomic orbitals and molecular orbitals
  - d. Give the chemical formulae of the coordination compounds
    - i. Potassium hexacyano ferrate (III)
    - ii. Trichloro triamine cobalt (III)
  - e. Calculate CFSE for d<sup>4</sup> system both strong and weak field?
  - f. How many microstates are possible for a d<sup>2</sup> configuration including both weak and strong field?
  - g. What do you mean orgal diagram?
  - h. Plot a curve between paramagnetic susceptibility  $(\chi_p)$  and diamagnetic susceptibility  $(\chi_D)$  with variation of temperature
  - i. Calculate the magnetic moments of [Cr(NH<sub>3</sub>)<sub>6</sub>]Br<sub>3</sub> and [Co(NH<sub>3</sub>)<sub>6</sub>]Cl<sub>3</sub>
  - j. How does nuclear reactions differ from chemical reactions?

PART - B (6 x 5 = 30 Marks)

#### Answer ANY FIVE questions

Marks

- 2. Show that oxygen molecule is paramagnetic based on molecular orbital theory (6)
- 3. What is meant by hybridisation of atomic orbitals? Draw the shapes of sp, sp<sup>2</sup> and sp<sup>3</sup> (6) hybrid orbitals
- 4. Explain the d-orbitals splitting diagrams of trigonal bi-pyramidal and square pyramidal complexes of  $ML_5$
- 5. Draw MO diagram for  $[Ti(H_2O)_6]^{3+}$  octahedral complex. Explain the formation of bonding, non-bonding and antibonding MOs and comment on spectra & its magnetic property
- 6. Discuss Gouy's method for measuring magnetic susceptibility of the complex (6)
- 7. Discuss the K<sub>2</sub>CrO<sub>4</sub>, KMnO<sub>4</sub>, HgI<sub>2</sub> and FeCl<sub>2</sub> compounds are coloured, why? (6)
- 8. a. How many  $\alpha$  -particles and  $\beta$ -particles are emitted in the decay of  $^{238}$ U to  $^{206}$ Pb? 3 + 3
  - b. What do you mean radioactive equilibrium?

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